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5 Reactions

Review exercise 5.1

1 a $2\text{Ca}(\text{NO}_3)_2(\text{s}) \rightarrow 2\text{CaO}(\text{s}) + 4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$
b $2\text{CuFeS}_2(\text{s}) + 4\text{O}_2(\text{g}) \rightarrow \text{Cu}_2\text{S}(\text{s}) + 2\text{FeO}(\text{s}) + 2\text{SO}_2(\text{g})$
c $2\text{KNO}_3(\text{s}) + 10\text{KI}(\text{s}) \rightarrow 6\text{K}_2\text{O}(\text{s}) + \text{N}_2(\text{g})$
d $4\text{NH}_3(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{N}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

2 a A pale yellow-green gas is bubbled through a colourless solution and the solution changes to an orange-red colour.
b A silver-gray solid is added to a blue solution. The solid becomes coated with a dark red solid and the blue colour of the solution slowly fades to colourless.

3 a $\text{Al}_2\text{O}_3(\text{s}) + 6\text{H}^+(\text{aq}) \rightarrow 2\text{Al}^{3+}(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$
A white solid dissolves when a colourless solution is added. A colourless solution remains after the reaction.
b $\text{H}_2\text{S}(\text{g}) + \text{Cl}_2(\text{aq}) \rightarrow 2\text{H}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{S}(\text{s})$
Colourless gas, with a rotten egg smell, bubbled through a colourless (or pale green-yellow) solution produces a yellow solid and a colourless solution. The rotten egg smell may disappear.
c $\text{CH}_3\text{COOH}(\text{aq}) + \text{HCO}_3^-(\text{aq}) \rightarrow \text{CH}_3\text{COO}^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
Two colourless solutions combine to produce a colourless, odourless gas and a colourless solution.
d $3\text{Cu}(\text{s}) + 8\text{H}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow 3\text{Cu}^{2+}(\text{aq}) + 2\text{NO}(\text{g}) + 4\text{H}_2\text{O}(\text{l})$
A colourless solution is added to a salmon-red solid and produces a colourless gas and a blue solution.

Review exercise 5.2

1 a silver iodide, AgI
b calcium phosphate, $\text{Ca}_3(\text{PO}_4)_2$
c no precipitate

2 a no reaction occurs
b $\text{Cu}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq}) \rightarrow \text{Cu}(\text{OH})_2(\text{s})$
When a blue solution and a colourless solution are mixed, a blue precipitate forms.
c $\text{Co}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CoCO}_3(\text{s})$
A pink solution and a colourless solution are mixed to form a pink precipitate.

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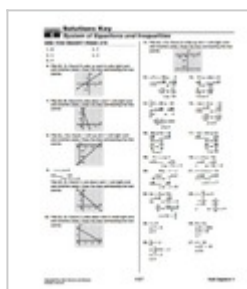
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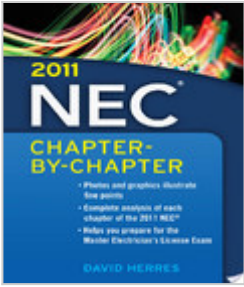
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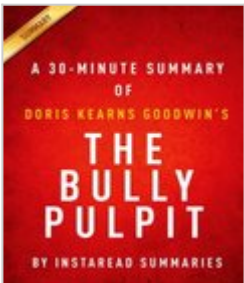
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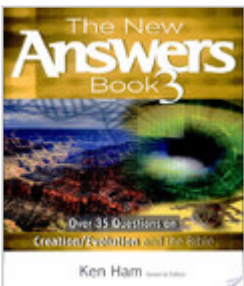
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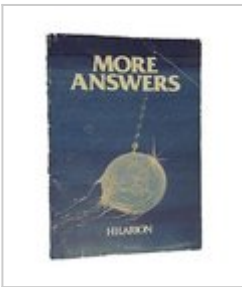
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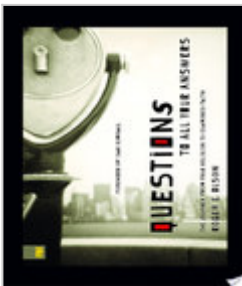
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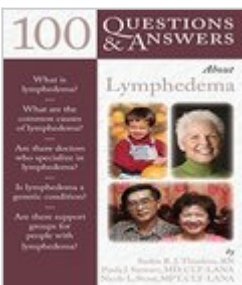
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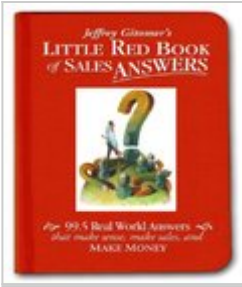
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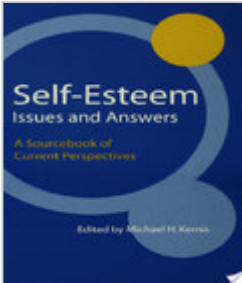
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